

Correction to Reactions of Metals in Nitric Acid: Writing Equations and Calculating Electromotive Force of Redox Reaction

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In the letter “Reactions of Metals in Nitric Acid: Writing Equations and Calculating Electromotive Force of Redox Reaction”,¹ the second sentence in the second paragraph should be corrected to read:

“In the present work, metals are divided into two groups on the basis of the value of their standard potential E° ^{5,6} in a typical activity series: reactive metals with $E^{\circ} < 0$ and less reactive ones, with $E^{\circ} > 0$.”

In addition, the left side of Figure 1A should read “Reactive $E^{\circ} < 0$ ” (instead of “Reactive $E^{\circ} > 0$ ”) and the left side of Figure 1B should be “Less reactive $E^{\circ} > 0$; Reactive $E^{\circ} < 0$ ” (instead of “Less reactive $E^{\circ} < 0$; Reactive $E^{\circ} > 0$ ”). The corrected figure is shown in Figure 1.

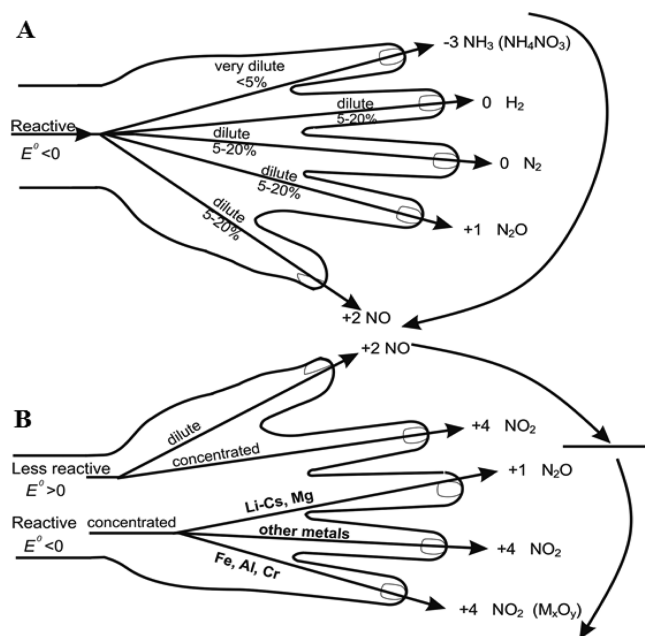


Figure 1. Suggested mnemonic schemes indicate the reduction product in nitric acid. (A) Left-hand schemes for reactive metals in dilute nitric acid; (B) right-hand scheme for less reactive metals in dilute and concentrated nitric acid, and for reactive metals in concentrated nitric acid.

REFERENCES

- (1) Sulcius, A. Reactions of Metals in Nitric Acid: Writing Equations and Calculating Electromotive Force of Redox Reaction. *J. Chem. Educ.* 2015, 92 (12), 1971–1972.